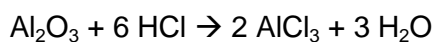


Question 1- Calculate the mass of 23 mole of

- Sulfuric acid H_2SO_4
- Aluminium hydroxide $\text{Al}(\text{OH})_3$
- Hydrated sodium thiosulfate $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$
- Potassium permanganate KMnO_4
- Potassium dichromate $\text{K}_2\text{Cr}_2\text{O}_7$
- Propanoic acid $\text{C}_3\text{H}_6\text{O}_2$
- Lead nitrate $\text{Pb}(\text{NO}_3)_2$
- Tetrachloromethane CCl_4

Question 2- Complete the table below



Mass of Al_2O_3	Amount of Al_2O_3	Volume of 0.1 mol L^{-1} HCl	Amount of HCl	Mass of AlCl_3	Amount of AlCl_3	Mass of H_2O	Amount of H_2O
35.0 g							
		13.2 mL					
				16.3 g			
						280 g	