

Question 1

- a) What mass of silver nitrate (AgNO_3) is required to make 250 mL of 0.100 molL^{-1} solution?
- b) What mass of sodium chloride (NaCl) is required to make 1L of 0.035 molL^{-1} solution?

Question 2

Calculate the concentration

- a) 2.31 g of sodium carbonate in water and diluting to 250mL
- b) 46.2 g of sodium hydroxide in water and diluting to 2L
- c) 1.22 g of sodium hydrogencarbonate in water and diluting to 200 mL
- d) 2.52 g of copper sulphate ($\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$) in water and diluting to 100 mL

Question 3

Copy and complete the following table

Concentration (molL^{-1})	Volume	Amount (mol)
0.1		2.57×10^{-3}
0.053	250 mL	
	3 L	2.732
0.457		0.256
0.107	20.7 mL	
	0.0132 L	5.27×10^{-4}